INTRODUCTION

The environmental contamination with heavy metals is causing an increasing concern nowadays due to the industrial growth, agricultural practices, and transportation means. These metals include lead (Pb), chromium (Cr), cadmium (Cd), mercury (Hg), copper (Cu), nickel (Ni) and zinc (Zn). Cadmium, copper and nickel are among the major contaminants of water sources with heavy metals [Amarioarei et al., 2017]. Improper disposal of heavy metals into water bodies, even at low concentrations, has become a major challenge due to their noxious impact on all creatures, persistence in nature and difficulty in being transformed by microorganisms; therefore, they accumulate in the ecosystem [Miretzky et al., 2004]. Conventional remediation options, such as ion exchange, biosorption [Hasan et al., 2010; Hasan et al., 2016; Elham et al., 2018], chemical precipitation, electrochemical, adsorption [Mohammad et al., 2013; Abidin et al., 2020], and membrane treatment [Al-Hemiri et al., 2012; Abed, 2014] have been used for the removal of heavy metals. These techniques are expensive, have logistical problems and technical complexity.

Constructed wetlands (CWs) are an alternative treatment method with special merits of being sustained, efficient, economically and
environmentally friendly treatment that mimics the natural processes [Stefanakis, 2019; Said et al., 2020]. Researchers used different aquatic and terrestrial plants in their pilot or field studies to treat different contaminants by various mechanisms such as containment or contaminants removal into the upper plants parts or turn them into less harmful forms [Ismail et al., 2020]. CWs are comprised of chemical, physical, and biological processes that influence each other, resulting in properly treated water. Many plants have high potential for successful uptake of contaminants. They have the ability to absorb and accumulate metals which are essential for their growth (Cu, Ni, Fe, Mn, Zn, and Mg) as well as non-essential metals (Cd, Cr, Pb, Co, Ag, Se, Hg).

In this study, a pilot scale horizontal sub-surface flow constructed wetlands (HSSF-CWs) were investigated to remove Cd\(^{2+}\), Cu\(^{2+}\), and Ni\(^{2+}\) ions from simulated wastewater. A Nerium oleander plant was used in this study due to its potential to remove heavy metals. Oleander is a native ornamental plant that is found in almost different Iraqi regions. It is fast-growing, evergreen, has high biomass, withstands extreme temperature and is easily transplanted into most soils, although it grows well in poor, resistant to salt stress [Kumar et al., 2017]. This is an ever-green, non-edible, terrestrial plant which can endure the Iraqi climate. The main objectives of the study were to examine the performance of the HSSF-CW system in removing Cd\(^{2+}\), Cu\(^{2+}\), and Ni\(^{2+}\) from simulated water under the climatic condition of Iraq and to determine the effect of substrate in heavy metals retention.

**MATERIALS AND METHODS**

**Simulation of heavy metal contaminated wastewater**

Cadmium nitrate Cd(NO\(_3\))\(_2\).4H\(_2\)O (Tetenal photowerk GmbH, Germany), copper nitrate Cu(NO\(_3\))\(_2\).3H\(_2\)O (BDH chemicals Ltd, UK), and nickel chloride Ni(Cl)\(_2\).6H\(_2\)O (Tetenal photowerk GmbH, Germany), all with purities above 98% were to prepare simulated heavy metal contaminated wastewater by dissolving them with tap water. The concentrations of Cd\(^{2+}\), Cu\(^{2+}\), and Ni\(^{2+}\) ions were calculated and adjusted at pH value of 5 to maintain the soluble cation forms.

**Pilot-Scale Horizontal Sub-Surface Flow - Constructed Wetland**

In this study, four planted glass basins were constructed outdoor and operated continuously with a dimension of 70×30×30 cm (LxWxD) and two different hydraulic retention times (HRTs) of 4 and 7 days were examined. A slope of (1%) was used to maintain the hydraulic gradient and aspect ratio of 2.3:1. Figures 1 and 2 show the photo and schematic diagram of HSSF-CW, respectively. Four storage tanks (each with 30 L capacity) were connected to four control tanks (each with 8 L capacity) with an opening to drain the excess simulated water that was fed from storage tanks to obtain constant water elevation throughout the whole study. The outlet valve of each control tank was attached to a stopcock to control the water flow into each basin. The flow rates were measured using a stopwatch and measuring cylinders. Each HSSF-CW
was connected to a PVC pipe which controls water level within each bed as well as for periodic collection of samples from the effluent, and a valve to empty the basins. Coarse gravel with a size of > 4 mm was added at 5 cm length from the inlet and outlet zones with depth of 20 cm in each basin in order to obtain a uniform distributed flow. Then, the remaining area of the basins was filled with 7 cm sand, 7 cm fine gravel and 6 cm coarse gravel. The oleanders used in this study had an average height of 25 cm and a root length ranging from 10–12 cm.

Hydraulic retention time (HRT) is the actual volume of wastewater in the bed divided by the average design flow. Many studies found that HRT is not sufficient to achieve the required contaminants removal [EPA, 1993]. Many researchers suggested that it is more convenient to apply a nominal detention time (\(\tau\)) as an alternative to HRT [Kadlec and Wallace, 2009]; [Rahi, 2019]. For all reactors, as the volume is known, nominal detention time (\(\tau\)) was calculated according to the formula given in equation (1):\[
\tau = \frac{V}{Q}
\]

where: \(\tau\) - hydraulic retention time (days); \(V\) - volume of the constructed wetland (m\(^3\)); \(Q\) - flow rate (m\(^3\)/d).

The total quantity of simulated wastewater used during each experiment was 45 L and was fed manually into each storage tank. Each basin was treating a volume of 5 and 9 L/d for 4- and 7-days HRT, respectively. The three CW were irrigated with \(\text{Cd}^{2+}\), \(\text{Cu}^{2+}\) and \(\text{Ni}^{2+}\), respectively, while the fourth basin was irrigated with tap water as a control. Treated heavy metal contaminated wastewater was sampled and analyzed.

**Analytical Methods**

**Heavy Metals Measurement in Simulated Wastewater**

The wastewater samples were analyzed daily. The samples were filtered using (Whatman No. 45) filter paper for the estimation of heavy metals, collected in plastic test tube and then analyzed using atomic absorption spectrometry (ASS) (SenSAA GBC Scientific Equipment, Malaysia). The percentage of \(\text{Cd}^{2+}\), \(\text{Cu}^{2+}\) and \(\text{Ni}^{2+}\) removal on each sampling day was determined with Eq. (2):

\[
\text{Removal} = \left(\frac{C_o - C_i}{C_o}\right) \times 100
\]

where: \(C_o\) - concentration of heavy metal entering the HSSF-CW (mg/l); \(C_i\) - concentration of heavy metal flowing out of the HSSF-CW (mg/l).

**Heavy Metals Measurement in Sand Samples**

The sand was sampled at the beginning and at the end of each experiment during the first and last day of the experiments. These samples were taken at different points of 5 cm depths and mixed to ensure representative sampling. After that, the samples were air-dried and digested. The total (\(\text{Cd}^{2+}\), \(\text{Cu}^{2+}\), \(\text{Ni}^{2+}\)) concentrations in sand was extracted using the hot aqua regia digestion procedure (ISO standard 11466). Then, the concentration of heavy metals was analyzed using ASS.

**Heavy Metals Measurement in Plant Tissues**

The plant tissues were analyzed on the first and last days of the experiments. Plant samples
RESULTS AND DISCUSSION

Heavy Metals Removal from Water

The results of the Cd\(^{2+}\) concentration during the 4 and 7 days HRT are shown in Figure 3 (a) and (b), respectively. At 4 days HRT, the Cd\(^{2+}\) concentration in the effluent ranged from 0.213 to 0.626 mg/l, which exceeds the permissible limits for local river conservation criteria (< 0.01 mg/l). The average removal efficiency was 96.5%. At 7 days HRT experiment, the Cd\(^{2+}\) concentrations in the effluent were 1.59 mg/l at day-1, 0.17 at day-5 sampling, and decreased to meet the permissible limits starting from day-7 sampling to be 0.008 mg/l. The average removal efficiency was 99.3% which was higher than the removal efficiency at 4 days HRT.

The results of the Cu\(^{2+}\) concentration at 4 and 7 days HRT in the effluent are shown in Figure 4 (a) and (b), respectively. During the 4 days HRT, the Cu\(^{2+}\) concentration at day 1 sampling was 3.63 mg/l, ending by 0.661 mg/l, which exceeds the permissible limits for local river conservation criteria (< 0.2 mg/l). The average removal efficiency

Chemical Parameters

The parameters of pH, dissolved oxygen (DO), and total dissolved solids (TDS) were measured to observe the potential changes in water. The pH in the water samples was measured by means of a HANNA pH meter model (HI 98103), TDS was measured using a HANNA conductivity tester model (HI 98301), and DO was measured using a HACH meter (HQ430d, USA).
was 96.3%. At 7 days HRT, the Cu\textsuperscript{2+} concentrations were 1.01 mg/l at day-1, 0.3 mg/l at day-5, and decreased to meet the permissible limits starting from day-6 with 0.088 mg/l. The average removal efficiency was 99.5%.

The results of Ni\textsuperscript{2+} concentration at 4 and 7 days HRT in the effluent are shown in Figure 5 (a) and (b), respectively. At 4 days HRT, the concentrations of Ni\textsuperscript{2+} exceeded the permissible limits for local river conservation criteria (< 0.2 mg/l) and ranged from 15.1–20.5 mg/l. The average removal efficiency was 71%. At 7 days HRT, the Ni concentration ranged of 7.6 to 9.41 mg/l, and also exceeded the permissible limits with an average removal efficiency of 86.3%.

Regression Models for Cd\textsuperscript{2+}, Cu\textsuperscript{2+}, and Ni\textsuperscript{2+} Removal in HSS-CW

The results of the regression analysis between the time (t) and effluent concentration (C) of the three metals are shown in Table 1. A square polynomial was used to calculate the concentrations of Cd\textsuperscript{2+}, Cu\textsuperscript{2+}, and Ni\textsuperscript{2+} ions in the effluent water from the CW. The analysis showed a significant relationship between the variables t and C (R\textsuperscript{2} > 0.9), and the quadratic equation can be used to predict the Cd\textsuperscript{2+}, Cu\textsuperscript{2+}, and Ni\textsuperscript{2+} concentration in the effluents of simulated wastewater by time.

**Heavy Metals in Sand**

The results of sand analyses are summarized in Table 2. The heavy metals were high on the last day for both experiments compared to the first day. This could be attributed to the gypsiferous and calcareous characteristics of the sand, in addition to the neutral pH, the presence of clay content and – to a lesser extent – organic matter [Pappalardo et al., 2010]. Moreover, it was clearly shown that higher HRT resulted in higher heavy metals accumulation in soil [Gola et al., 2016].

**Heavy Metals Concentration in Plant Tissues**

The results of the plant tissues analyses are summarized in Table 3. The oleanders planted near

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**Table 1.** Regression of Cd\textsuperscript{2+}, Cu\textsuperscript{2+}, and Ni\textsuperscript{2+} concentration in the effluent with time

<table>
<thead>
<tr>
<th>Metals</th>
<th>HRT</th>
<th>Regression equation</th>
<th>R\textsuperscript{2}</th>
<th>F-value</th>
<th>P-value</th>
</tr>
</thead>
<tbody>
<tr>
<td>Cd\textsuperscript{2+}</td>
<td>4</td>
<td>C = 0.055t\textsuperscript{2} - 0.411t + 0.979</td>
<td>0.9844</td>
<td>63.0</td>
<td>&gt; 0.05</td>
</tr>
<tr>
<td>Cu\textsuperscript{2+}</td>
<td>7</td>
<td>C = 0.049t\textsuperscript{2} - 0.652t + 2.075</td>
<td>0.8978</td>
<td>26.4</td>
<td>&gt; 0.05</td>
</tr>
<tr>
<td>Ni\textsuperscript{2+}</td>
<td>4</td>
<td>C = 0.163t\textsuperscript{2} - 1.773t + 5.376</td>
<td>0.9749</td>
<td>38.8</td>
<td>&gt; 0.05</td>
</tr>
<tr>
<td>Ni\textsuperscript{2+}</td>
<td>7</td>
<td>C = 0.057t\textsuperscript{2} - 0.996t + 10.028</td>
<td>0.9292</td>
<td>21.0</td>
<td>&gt; 0.05</td>
</tr>
</tbody>
</table>

**Table 2.** Concentration of Cd\textsuperscript{2+}, Cu\textsuperscript{2+}, and Ni\textsuperscript{2+} in sand at 4 and 7 days HRT

<table>
<thead>
<tr>
<th>HRT (d)</th>
<th>Sampling time</th>
<th>Cd (mg/kg)</th>
<th>Cu(mg/kg)</th>
<th>Ni (mg/kg)</th>
</tr>
</thead>
<tbody>
<tr>
<td>4</td>
<td>1st day</td>
<td>8</td>
<td>21</td>
<td>30</td>
</tr>
<tr>
<td></td>
<td>Last day</td>
<td>21</td>
<td>10</td>
<td>48</td>
</tr>
<tr>
<td>7</td>
<td>1st day</td>
<td>7</td>
<td>17</td>
<td>41</td>
</tr>
<tr>
<td></td>
<td>Last day</td>
<td>44</td>
<td>63</td>
<td>287</td>
</tr>
</tbody>
</table>
the HSSF-CW inlet showed higher concentrations accumulated in their tissues than those planted near the outlet. This could be attributed to the metal ions absorbed by the plant along with their flow through the HSSF-CW, in addition to the metal sorption capacity of the sand. At 7 days HRT, the oleander plant accumulated higher Cd$^{2+}$, Cu$^{2+}$, and Ni$^{2+}$ than at 4 days HRT. Moreover, it was obvious that oleander accumulated more Cd$^{2+}$, Cu$^{2+}$, and Ni$^{2+}$ at the last day of the experiments than the first day.

### Monitoring of pH, DO and TDS

The pH, DO and TDS were recorded for each metal throughout the study and the profiles are shown in Figures 6, 7, and 8, respectively.

Throughout the 4 days HRT, the pH values ranged from 7.1 to 7.6 for Cd$^{2+}$, 6.7 to 7.5 for Cu$^{2+}$, 7.1 to 7.5 for Ni$^{2+}$, and 7.2 to 7.8 for control. At the 7 days HRT, the pH ranged from (7.5–7.8) for Cd$^{2+}$, (7.4–7.5) for Cu$^{2+}$, (7.4–7.6) for Ni$^{2+}$, (7.5–7.91) control. The results did not show any significance between the different CWs and both HRTs.

The DO ranged between 5.1 to 7.2 mg/l for Cd$^{2+}$, 4.3 to 6.7 mg/l for Cu$^{2+}$, 3.6 to 6.4 mg/l for Ni$^{2+}$, and 6.1 to 7.4 mg/l for control during the 4 days HRT. At 7 days HRT, the DO ranged between 3.4 to 6.2 mg/l for Cd$^{2+}$, 3.3 to 5.4 mg/l for Cu$^{2+}$, 4.1 to 5.8 mg/l for Ni$^{2+}$, and 4.3 to 7.2 mg/l for control. According to Podedworna and Zubbrowska-Sudol [2010], these

### Table 3. Heavy metals accumulation in oleander at the beginning and the end of CW with HRT of 4 and 7 days

<table>
<thead>
<tr>
<th>HRT (d)</th>
<th>Sampling time</th>
<th>Cd (mg/kg)</th>
<th>Cu (mg/kg)</th>
<th>Ni (mg/kg)</th>
</tr>
</thead>
<tbody>
<tr>
<td></td>
<td></td>
<td>Inlet</td>
<td>Outlet</td>
<td>Inlet</td>
</tr>
<tr>
<td>4</td>
<td>1st day</td>
<td>67</td>
<td>23</td>
<td>90</td>
</tr>
<tr>
<td></td>
<td>Last day</td>
<td>81</td>
<td>38</td>
<td>154</td>
</tr>
<tr>
<td>7</td>
<td>1st day</td>
<td>57</td>
<td>33</td>
<td>101</td>
</tr>
<tr>
<td></td>
<td>Last day</td>
<td>109</td>
<td>89</td>
<td>279</td>
</tr>
</tbody>
</table>

![Figure 6](image1.png)  
**Figure 6.** The pH values measured at the effluent of HSSFCW with HRT of a) 4 days b) 7 days

![Figure 7](image2.png)  
**Figure 7.** The DO values measured at the effluent of HSSFCW with HRT of a) 4 days b) 7 days
results indicated that all tests were conducted under aerobic conditions. The TDS values during the 4 days HRT experiment ranged from 974 to 1830 mg/l in the effluent of the HSSF-CWs irrigated with Cd$^{2+}$, Cu$^{2+}$ and Ni$^{2+}$ and 973 to 1174 mg/l for control. At 7 days HRT, the TDS values ranged from 1187 to 1781 mg/l for the CWs irrigated with Cd$^{2+}$, Cu$^{2+}$ and Ni$^{2+}$ and 1277 to 1396 mg/l in the control. In general, the TDS values were higher for HSSF-CWs treated with Cd$^{2+}$, Cu$^{2+}$, and Ni$^{2+}$. This result agrees with Borker et al., [2013] who found that TDS in water increased due to the metal ions added to water addition.

CONCLUSIONS

Higher removal efficiencies from simulated wastewater were observed in 7 days HRT with an average removal efficiency of 99.3, 99.5, and 86.3% for Cd$^{2+}$, Cu$^{2+}$, and Ni$^{2+}$, respectively. At 4 days HRT, the removal efficiencies were only 96.5 (Cd$^{2+}$), 96.3 (Cu$^{2+}$), and 71% (Ni$^{2+}$). This indicated the essential role of HRT on the treatment efficiency of heavy metals in HSSF-CWs. The HSSF-CW was able to reduce the Cd$^{2+}$ and Cu$^{2+}$ concentrations to meet the permissible limits at 7 days HRT, while increasing HRT has a slight contribution on the removal of Ni$^{2+}$ ions and it is obvious that the metal concentrations were higher than the permissible limits at both HRTs.

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REFERENCES


